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Combined In Situ XAFS/DRIFTS Studies of the Evolution of Nanoparticle Structures from Molecular Precursors

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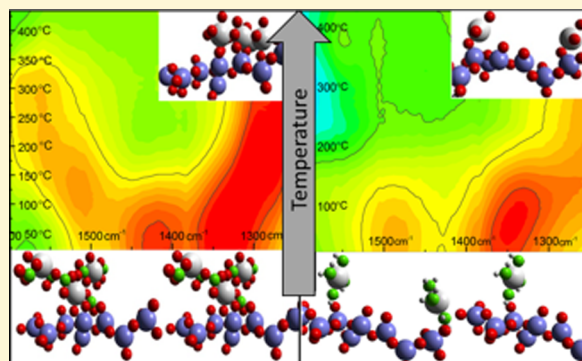
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Supporting Information

ABSTRACT: The rational design of catalyst materials is of great industrial significance, yet there is a fundamental lack of knowledge in some of the most well-established processes, e.g. formation of supported nanoparticle structures through impregnation. Here, the choice of precursor has a significant influence on the resulting catalytic properties of the end material, yet the chemistry that governs the transformation from defined molecular systems to dispersed nanoparticles is often overlooked. A spectroscopic method for advanced in situ characterization is employed to capture the formation of PdO nanoparticles supported on γ -Al₂O₃ from two alternative molecular precursors - Pd(NO₃)₂ and Pd(NH₃)₄(OH)₂. Time-resolved diffuse reflectance infrared Fourier transform spectroscopy is able to identify the temperature assisted pathway for ligand decomposition, showing that NH₃ ligands are oxidized to N₂O and NO⁻ species, whereas NO₃⁻ ligands assist in joining Pd centers via bidentate bridging coordination. Combining with simultaneous X-ray absorption fine structure spectroscopy, the resulting nucleation and growth mechanisms of the precious metal oxide nanoparticles are resolved. The bridging ability of palladium nitrate aids formation and growth of larger PdO nanoparticles at lower onset temperature (<250 °C). Conversely, impregnation from [Pd(NH₃)₄]²⁺ results in well-isolated Pd centers anchored to the support which requires a higher temperature (>360 °C) for migration to form observable Pd–Pd distances of PdO nanoparticles. These smaller nanoparticles have improved dispersion and an increased number of step and edge sites compared to those formed from the conventional Pd(NO₃)₂ salt, favoring a lower light off temperature for complete methane oxidation.



INTRODUCTION

Supported metal nanoparticles are a cornerstone of heterogeneous catalysis, and by extension, the chemical industry. Optimizing the preparation method for smaller nanoparticle size and improved dispersion is a common theme in heterogeneous catalyst design: to improve specific surface area, lower precious metal content, and increase metal–support interfacial regions while maintaining thermostability toward sintering. There are various methods for the preparation of supported metal catalysts such as controlled colloidal routes,^{1,2} deposition precipitation,³ grafting techniques,⁴ and atomic deposition, all of which result in nanoparticle materials with

differing properties. Despite these efforts, the conventional route of incipient wetness impregnation is regarded as a reliable route, taking advantage of accessible precursor materials (inorganic metal salt) in a one-step batch process that can be scaled up for industrial production.⁵

Adjustments to the impregnation method by using alternative metal precursors result in differing properties of the final nanoparticle catalyst together with any changes to

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temperature and atmospheric conditions of the proceeding drying and calcination steps.^{6,7} The amphoteric nature of common metal oxide supports means that the pH of the impregnating solution is an important factor. Charged surface sites (e.g., Al–O[−] or Al–OH⁺) can be created when the pH is above or below the isoelectric point of the support, which allow the impregnation of metal salts to occur via ion exchange or ligand substitution mechanisms.⁸ Nanoparticles prepared from chloride containing precursors are found to suffer from residual Cl[−] blocking surface catalytic sites,^{9–11} whereas organic precursors such as acetate or acetylacetonate as well as ammonia containing salts and nitrate precursors have been reported to yield greater dispersion over the support.^{6,9,12} Despite investigations into the influence of preparation methods on the final metal dispersion and catalytic activity,^{7,13} there is often lack of evidence and reasoning for the chemical processes that govern the nanoparticle formation.

In this study, genesis of PdO nanoparticles from two different impregnated Pd precursors, Pd(NO₃)₂ and Pd(NH₃)₄(OH)₂, are investigated. The development of PdO nanoparticles on high surface area γ -Al₂O₃ has been studied extensively for a number of catalytic applications: CO oxidation,¹⁴ H₂O₂ synthesis,¹⁵ upgrading bio-oils,¹⁶ and automotive three-way catalysts.¹⁷ Of recent industrial interest is the use of Pd/ γ -Al₂O₃ in catalyzing low temperature methane oxidation for the clean and efficient operation of compressed natural gas engines (CNGs).^{18–21} Mechanistic studies propose the oxidation of CH₄ to occur on a reducible PdO nanoparticle surface, close to the Al₂O₃ interface where oxygen migration from the support can assist in reoxidation of the Pd nanoparticle surface via a Mars–Van Krevelen mechanism.²¹ Therefore, the preparation of Pd/ γ -Al₂O₃ with greatest dispersion of PdO for increased interfacial contact with the support is most desirable.

The advantage of synchrotron-based techniques applied to investigate industrial catalyst preparation is highlighted in work by Espinosa-Alonso et al. for the formation of supported Ni materials.²² In this study, by combining two noninvasive, time-resolved spectroscopic techniques, the surface and bulk properties of the precursor materials are probed in situ under a controlled calcination environment. While diffuse reflectance infrared Fourier transform spectroscopy (DRIFTS) is used to follow the vibrational modes of adsorbed and coordinated inorganic molecular species at the surface, X-ray absorption fine structure (XAFS) is used to follow the local coordination environment and oxidation state of Pd throughout the bulk of the sample. Using both techniques at the same time with online mass spectrometry of the effluent gas gives the advantage of confidently assigning features that change with response to sample environment and allow interpretation of the mechanisms involved in ligand decomposition and metal nanoparticle formation.

EXPERIMENTAL SECTION

Sample Preparation. Two 3 wt % Pd/ γ -Al₂O₃ catalysts were prepared by incipient wetness impregnation of an aqueous solution of Pd salt onto a γ -Al₂O₃ support (SASOL), followed by calcination (500 °C, 2 h). The first sample was prepared from an acidified aqueous solution of palladium nitrate (15.11 wt % Pd, Johnson Matthey) and a defined quantity of HNO₃, corresponding to a molecular ratio NO₃[−]/Pd = 3.8. The second sample was prepared from an aqueous solution of tetraamminepalladium(II) hydroxide (5.96 wt % Pd, Johnson Matthey), prepared by dissolution of tetraamminepalladium hydrogen carbonate (Johnson Matthey, 99.99%) in aqueous NH₄OH solution

(28 wt %). Incipient wetness impregnations were realized at room temperature with calculated quantities of respective aqueous Pd solutions to obtain 3 wt % Pd catalysts. The impregnated precursor catalyst samples Pd(NO₃)₂/ γ -Al₂O₃ and Pd(NH₃)₄(OH)₂/ γ -Al₂O₃ were subsequently dried at 100 °C overnight.

In Situ XAFS/DRIFTS Measurements. In situ XAFS and DRIFTS measurements were conducted at B18, Diamond Light Source, Didcot, UK of the impregnated, dried precursor samples during the calcination procedure using the modified Harrick DRIFTS cell. A DaVinci arm fitted with praying mantis optics was used to refocus the IR beam outside the FTIR spectrometer for positioning of the Harrick DRIFTS cell in the X-ray beam. The Harrick XAFS/DRIFTS cell, which has been used and reported previously,²³ has an X-ray path length of 3.17 mm positioned 1.04 mm below the surface of the catalyst bed. The sample was heated to 500 °C (10 °C min^{−1} under a constant controlled flow of air (30 mL min^{−1}) and maintained at 500 °C for a 2-h dwell period before being cooled back to room temperature. XAFS and DRIFTS spectra were recorded continuously throughout this period, and the effluent gas was monitored online by a Hiden QGA mass spectrometer. XAFS measurements were performed at the Pd K-edge in transmission mode using QEXAFS setup with a fast scanning Si(311) double crystal monochromator. All XAFS spectra were acquired with a Pd foil placed between I_t and I_{ref} and the time resolution was 21 s spectrum^{−1} (k_{max} = 15.9). DRIFTS spectra were collected with an Agilent Carey 680 FTIR spectrometer taking 64 scans with a resolution of 4 cm^{−1} using the liquid nitrogen cooled MCT detector. The time resolution of DRIFTS was 30 s spectrum^{−1}.

EXAFS Analysis. Processing extended X-ray absorption fine structure (EXAFS) data was performed using IFEFFIT with the Horae package (Athena and Artemis). Athena was used to calibrate, align, and normalize the spectra with respect to the Pd foil, for which E_0 was set at 24 358 eV. EXAFS data processing of k^2 data used an appropriate k range for the data (3.2–13.4 Å^{−1}). Analysis of the Fourier transformed data was limited to fitting with the first two coordination shells (denoted Pd_{(O)1}, Pd_{(O)2}, Pd_{(N)1}, Pd_{(N)2}, Pd_{(Pd)1}, and Pd_{(Pd)2}), using cif files of Pd(NO₃)₂ and PdO. The amplitude reduction factor, S_0^2 , was derived from fitting the Pd foil using a coordination number of 12 to give a value of 0.8. The value of σ^2 , the mean square relative displacement of absorber and backscatter atoms, is known to increase with temperature.²⁴ Consequently, the linear dependence of σ^2 with temperature was fitted using nonphase corrected Fourier transformed Pd K-edge data collected of the calcined PdO/ γ -Al₂O₃ between 500 to 20 °C, where all other parameters (Δr and S_0^2) are assumed to be fixed. This variance in σ^2 with temperature, reported in Figure S1 of the [Supporting Information](#), was used to fix σ^2 for the fitting of the EXAFS data collected during the calcination temperature ramp period.

Ex Situ Characterization. Thermogravimetric analysis was carried out in a TA thermogravimetric analyzer (instrument Q50/DSC Auto Q20) with a sample mass of 10 mg placed in a platinum crucible under continuous flow of 40% N₂ and 60% air and ramped at 10 °C min^{−1}. Transmission electron microscopy (TEM) images of the samples were obtained using a JEOL JEM 2100 transmission electron microscope. Samples were prepared for TEM analysis by dispersing in high-purity ethanol using ultrasonication. Ten microliters of the sonicated suspension were pipetted onto a holey carbon supporting Cu grid, and the solvent evaporated. Ex situ Pd K-edge XAFS spectra of the calcined catalyst samples after reducing in H₂ (at 100 °C) were taken in transmission mode on B18 at Diamond Light Source. CO chemisorption of the calcined catalyst samples was performed with FTIR analysis in transmission mode. Samples (20 mg) were pressed into self-supporting wafers (1.35 cm²), mounted in the transmission FTIR cell, and dried at 100 °C under constant flow of He. Several pulses of CO were introduced into the cell until saturation. The CO adsorption FTIR spectra were recorded (40–4500 cm^{−1}, 4 cm^{−1} resolution) after purging the cell with He (15 min). Raman spectra were recorded using a Bruker Senterra microscope at B22, Diamond Light Source, using 532 nm laser at 5 mW. XRD diffraction patterns were collected using a Rigaku Miniflex diffractometer (ISIS Materials Characterization Lab) with a 600 mW Cu α tube source operated at 45

kV and 15 mA and using a 1D Si strip detector. Sherrer calculations of the average crystallite sizes used the full width half-maximum of the diffraction peaks located at the 2θ angle 55° , assuming the Sherrer constant to be 0.94. Catalytic light off curves for methane oxidation were obtained using JM automotive test rig, using a synthetic natural gas mixture (0.5% CH_4 , 10% O_2 , 0.01% C_2H_6 , 0.01% C_3H_8 , 0.01% NO , 10% CO_2 , and 10% H_2O), a gas hourly space velocity of 90 000 h^{-1} , and a temperature ramp rate of 15 $^\circ\text{C min}^{-1}$. Effluent gases were analyzed by a MKS 2000 multigas FTIR analyzer.

RESULTS AND DISCUSSION

Ex Situ Characterization. Two 3 wt % PdO catalysts supported on $\gamma\text{-Al}_2\text{O}_3$ were prepared by the impregnation and calcination from $\text{Pd}(\text{NO}_3)_2$ and $\text{Pd}(\text{NH}_3)_4(\text{OH})_2$ to give catalysts $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NO}_3)$ and $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NH}_3)$, respectively. TEM images of the calcined catalysts (Figure 1)

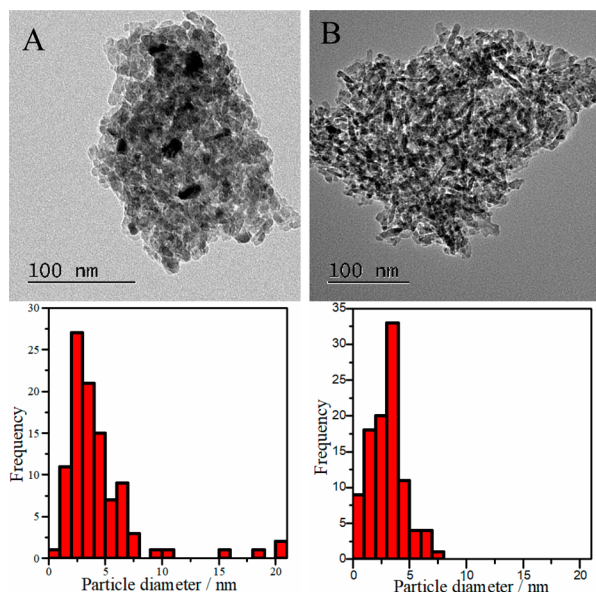


Figure 1. Particle size histogram and representative image of (A) $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NO}_3)$ and (B) $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NH}_3)$.

show dark Pd nanoparticles dispersed upon the contrasting lighter colored $\gamma\text{-Al}_2\text{O}_3$ support. Particle size analysis conducted over 100 measured particles revealed smaller average particle diameter (3.1 nm) for catalyst $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NH}_3)$ compared to that of $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NO}_3)$ (4.4 nm). This is consistent with the presence of larger agglomerated particles (>15 nm) observed only for $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NO}_3)$, as shown clearly in the corresponding histogram of particle size distribution. Although only 5% of the 100 measured particles have a diameter greater than 15 nm, this is significant when considering the mass fraction of Pd located at the core of larger particles rather than contributing to catalysis at the surface.

X-ray diffraction patterns of both calcined catalyst samples show diffraction peaks assigned to cubic $\gamma\text{-Al}_2\text{O}_3$ (JPDs card no. 00-002-1420) similar to that of the fresh support, labeled with arrows in Figure 2, proving that structural integrity of the support is maintained after impregnation. Additional diffraction peaks located at 2θ angles of 42.6 , 55.6 , and 72.7° , labeled with asterisks in Figure 2, can be attributed to crystalline domains of tetragonal PdO (JCPDS card no. 00-002-1432). The full width half-maximum of these peaks indicates the relative crystallite size of the observed PdO particles pursuant to the Sherrer equation.²⁵ Smaller intensity and increased broadening of PdO

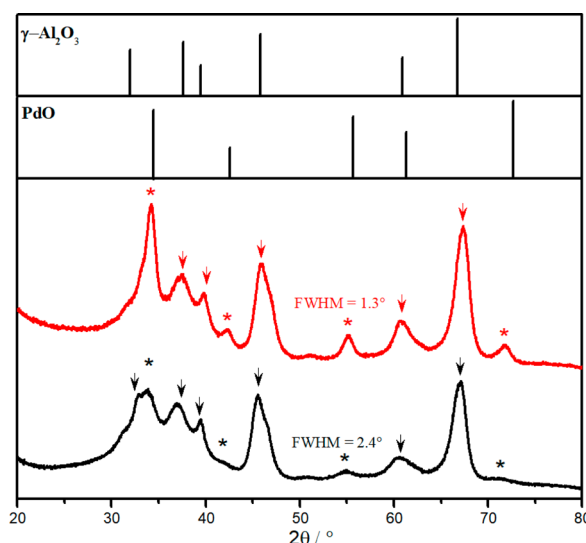


Figure 2. XRD patterns of calcined catalysts $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NO}_3)$ and $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NH}_3)$, shown in red and black, respectively. Top panels show the relative position and intensity of diffraction lines of $\gamma\text{-Al}_2\text{O}_3$ and PdO references.

diffraction peaks for $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NH}_3)$ suggests this sample has smaller PdO crystallite size of 6 nm compared to 11 nm for $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NO}_3)$. Further evidence for the PdO structure is indicated by the Raman spectra of $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NO}_3)$ and $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NH}_3)$ (Supporting Information, Figure S2), which show well-resolved bands positioned at 439 and 640 cm^{-1} associated with Raman active E_g and B_{1g} vibrational modes of Pd–O, respectively, in a PdO lattice. Difference in the respective peak area of the B_{1g} band (averaged from 6 spectra taken at different positions over each sample) indicates a greater number of Pd–O vibrations from $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NO}_3)$ compared to $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NH}_3)$, again indicating larger PdO crystallite size for $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NO}_3)$.²⁶

Ex situ Pd K-edge XAFS spectra of the calcined catalyst samples were collected for compositional and quantitative information on the Pd environment in each case. The Pd K-edge XANES show the Pd environment of each sample to be broadly consistent with PdO at room temperature (Supporting Information, Figure S3) with small differences in amplitude attributed to particle size effects. Under an H_2 atmosphere, PdO nanoparticles of both samples were completely reduced to metallic Pd. A fitting model to the EXAFS data (Figure 3a) was achieved using the first shell scattering path, Pd–Pd, of metallic Pd with σ^2 value of 0.0065 \AA^2 . The amplitude of this fitted scattering feature gave Pd–Pd coordination numbers of 10.3(4) for $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NO}_3)$ and 7.5(2) for $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NH}_3)$, corresponding to average particle diameters of 2.7(6) and 1.1(1) nm, respectively, using a previously reported method detailed in the Supporting Information (Tables S1 and S2).²⁷

CO can be used as a probe molecule in FTIR adsorption experiments to investigate surface adsorption properties of the catalyst materials.^{28,29} Linear and bridge bonding of CO adsorbed on both $\text{Pd}/\gamma\text{-Al}_2\text{O}_3$ samples were identified by the sharp band at 2071 cm^{-1} and the broad band between 1969 and 1830 cm^{-1} , respectively (Figure 3b).^{30–33} The significant difference in intensity of CO absorption bands from the two catalyst samples is consistent with $\text{Pd}/\gamma\text{-Al}_2\text{O}_3\text{-ex}(\text{NH}_3)$ having larger surface area due to smaller particle size. Additionally, the type of surface sites can be clarified from the relative linear to

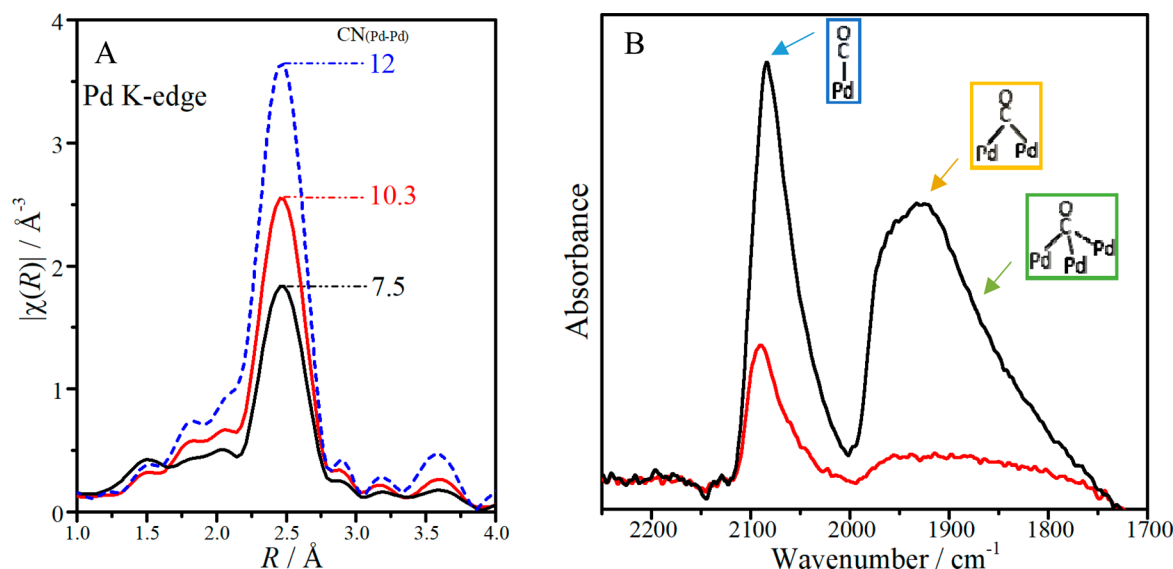


Figure 3. (A) Pd K-edge Fourier transform EXAFS data of calcined catalysts Pd/ γ -Al₂O₃-ex(NO₃) (red) and Pd/ γ -Al₂O₃-ex(NH₃) (black) reduced under H₂ atmosphere and Pd foil (blue). (B) Transmission FTIR spectra of calcined catalysts Pd/ γ -Al₂O₃-ex(NO₃) (red) and Pd/ γ -Al₂O₃-ex(NH₃) (black) after CO adsorption normalized to 20 mg catalyst.

bridge bonding adsorption band intensity ratio, reported in Table 1. The greater ratio of linear to bridge bonded CO on

Table 1. Calcined Pd/ γ -Al₂O₃ Catalyst Characterization^a

	average particle size (nm)		PdO crystallite size (XRD) (nm)	L/B ratio (CO adsorption FTIR)	T ₅₀ (°C)
	TEM	EXAFS			
Pd/ γ -Al ₂ O ₃ -ex(NO ₃)	4.4	2.7(6)	11	0.14	394
Pd/ γ -Al ₂ O ₃ -ex(NH ₃) ₄ (OH) ₂	3.1	1.1(1)	6	0.27	385

^aTEM average particle size, EXAFS average particle size, XRD PdO crystallite size, linear/bridged ratio (CO adsorption FTIR), and temperature required to achieve 50% CH₄ conversion during lean CH₄ oxidation catalytic test (T₅₀).

Pd/ γ -Al₂O₃-ex(NH₃) is indicative of a rough Pd surface with greater number of step and edge sites.^{32,54} The discrepancy between the average particle size calculated by the different methods, reported in Table 1, is not surprising. The TEM is limited to the resolution and field of view of the microscope; XRD is limited to assess particles with sufficiently large crystalline domains, while EXAFS probes all Pd species in the sample. Despite these differences, there is consistent evidence for Pd/ γ -Al₂O₃-ex(NH₃) providing smaller average PdO crystallite size and greater dispersion over the support compared to Pd/ γ -Al₂O₃-ex(NO₃). Light off curves for CH₄ conversion (Supporting Information, Figure S4) confirm that the catalyst with improved dispersion and smaller particle size, Pd/ γ -Al₂O₃-ex(NH₃), is able to catalyze the complete oxidation of methane at a temperature that is lower than that of the catalyst possessing some larger agglomerated particles, Pd/ γ -Al₂O₃-ex(NO₃).

In Situ XAFS/DRIFTS of the Preparation Route. Pd K-edge EXAFS analysis of the impregnated precursor samples, Pd(NO₃)₂/ γ -Al₂O₃ and Pd(NH₃)₄(OH)₂/ γ -Al₂O₃ gave structural information on the initial Pd coordination environment after impregnation onto the γ -Al₂O₃ at room temperature. The

k^2 weighted $\chi(k)$ EXAFS spectrum of the impregnate Pd(NH₃)₄(OH)₂/ γ -Al₂O₃ compares closely to that of the unsupported aqueous Pd(NH₃)₄(OH)_{2(aq)} solution (Figure 4a). Conversely, the k^2 weighted $\chi(k)$ EXAFS spectra of the impregnated precursor Pd(NO₃)₂/ γ -Al₂O₃ shows greater long-range structure from EXAFS oscillations extending to greater k range ($\sim 10.4 \text{ \AA}^{-1}$) with additional scattering feature at 9 \AA^{-1} (annotated with an asterisk in Figure 4b) that is absent from the $\chi(k)$ EXAFS spectrum of the unsupported Pd-(NO₃)₂(H₂O)_x reference. A simple fitting model was constructed from one calculated scattering path representative of four light atomic neighboring atoms ($14 \leq \text{atomic number}, Z \leq 16$) located at 2.049 \AA from the absorbing Pd atom. This model was sufficient to fit the Fourier transformed EXAFS data of unsupported Pd(NH₃)₄(OH)_{2(aq)} and impregnated Pd-(NH₃)₄(OH)₂/ γ -Al₂O₃ (Figure 4c), showing that the Pd²⁺ center in each case is coordinated to four oxygen or nitrogen atoms of ammonia, hydroxyl, or water ligands with little or no longer range structure, consistent with a [Pd(NH₃)₄]²⁺ complex. This model was, however, inadequate to fit that of the impregnated Pd(NO₃)₂/ γ -Al₂O₃. In this case, additional structure observed in the EXAFS data requires an additional scattering path to refine the EXAFS fitting model to this sample. A refined fit to Pd(NO₃)₂/ γ -Al₂O₃, displayed as the red dashed line of Figure 4d, includes an additional scattering contribution from neighboring Pd atoms located 3.43 \AA from absorbing Pd. The distance is consistent with a Pd atom in a second coordination shell from the absorbing Pd, which corresponds to neighboring Pd centers adjoined by bridging ligand coordination. It is noted that such a Pd distance (3.43 \AA) is close to that of the second coordination shell Pd atom in the crystal structure of tetrahedral PdO.

DRIFTS spectra of both precursor samples before calcination are reported in Figure 5a. The broad band centered at 3500 cm^{-1} can be assigned to symmetric and antisymmetric O–H stretching vibrations of lattice water, and a sharp band at 1420 cm^{-1} is assigned to carbonate ions on the oxide support.³⁵ Loss of these absorption bands indicate removal of H₂O and CO₃²⁻ upon heating to $200 \text{ }^\circ\text{C}$. The intensity of the FTIR absorption

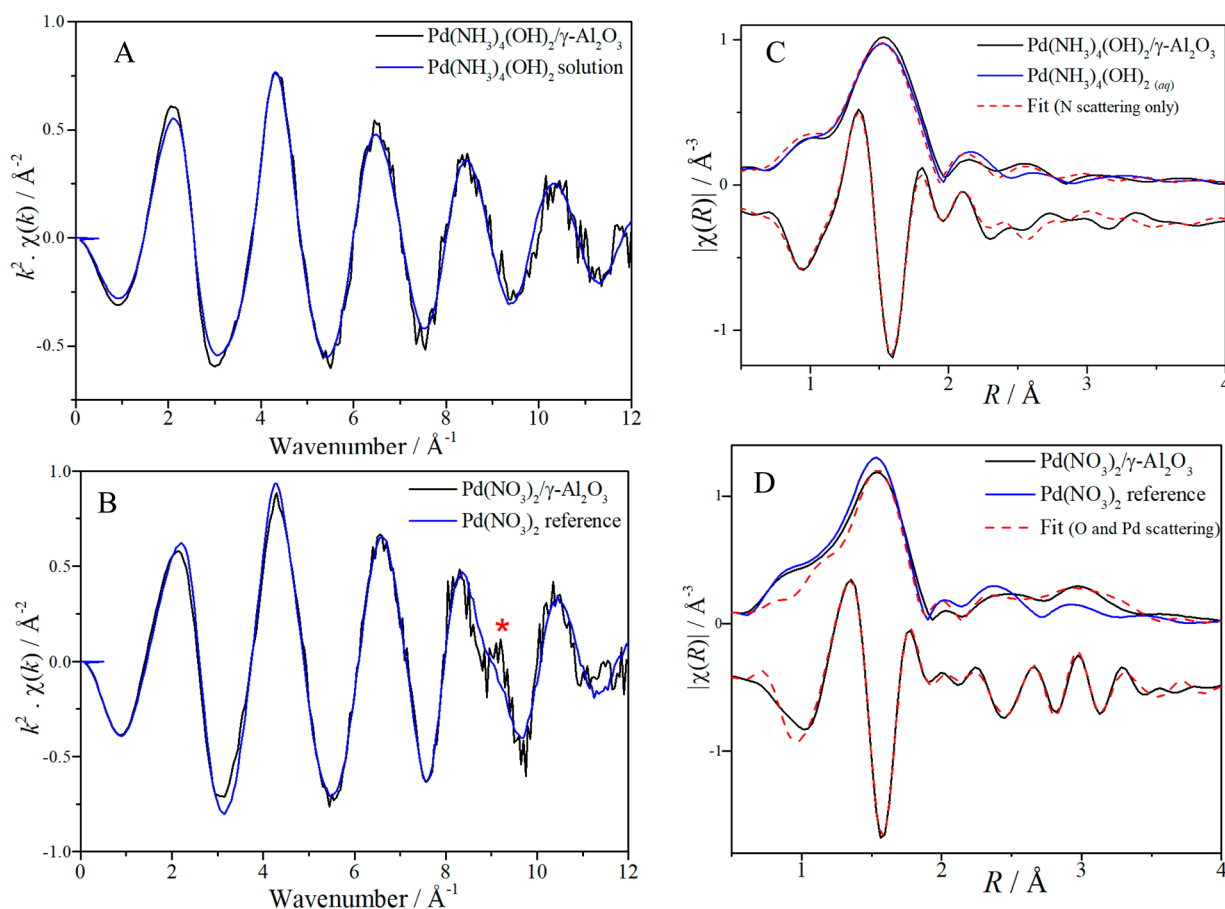


Figure 4. Pd K-edge k^2 weighted $\chi(k)$ EXAFS spectra and corresponding Fourier transform EXAFS data of $\text{Pd}(\text{NH}_3)_4(\text{OH})_2/\gamma\text{-Al}_2\text{O}_3$ (top) and $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$ (bottom). Nonphase corrected Fourier transformed EXAFS data of precursor sample (black) plotted with corresponding unsupported reference $\text{Pd}(\text{NH}_3)_4(\text{OH})_2$ (aq) and $\text{Pd}(\text{NO}_3)_2$ (blue) and the fitting model using first shell $\text{Pd}_{(\text{N})1}$ or $\text{Pd}_{(\text{O})2}$ and second shell $\text{Pd}_{(\text{Pd})2}$ scattering paths (red, dashed).

bands in the spectral range 1250–1800 cm^{-1} during the calcination temperature ramp are shown as a color map in Figure 5b. The nitrate species of sample $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$ can be identified by two molecular vibrational bands of N=O stretching between 1650 and 1500 cm^{-1} and O–N–O asymmetrical stretching between 1350 and 1200 cm^{-1} ,^{36,37} consistent with the calculated N=O stretching frequency of bridging nitrates on an Al_2O_3 surface.³⁸ Comparison with the DRIFTS spectra of the unsupported $\text{Pd}(\text{NO}_3)_2(\text{H}_2\text{O})_x$ precursor and an acid treated $\gamma\text{-Al}_2\text{O}_3/\text{HNO}_3$ (Supporting Information, Figure S5) shows the O–N–O asymmetrical stretching of nitrate coordinated to Pd^{2+} and $\gamma\text{-Al}_2\text{O}_3$ as absorption bands at 1320 and 1300 cm^{-1} , respectively. The broad character of the O–N–O stretching band of $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$ can be seen to result from nitrate coordinated to both the Pd^{2+} cation and Lewis acid Al^{3+} sites of the $\gamma\text{-Al}_2\text{O}_3$. Beyond 100 °C, the broad band initially centered at 1300 cm^{-1} begins to narrow and shift to lower frequency until it is centered at 1270 cm^{-1} by 200 °C. The nitrate band previously centered at 1510 cm^{-1} also broadens and shifts over the same temperature range but toward higher wavenumbers until it is centered at 1570 cm^{-1} by 200 °C (annotated by arrows, Figure 5b). The shift in nitrate bands to greater separation is consistent with increased bidentate nitrate coordination to a metal center.^{38,39} The early preassociation of Pd species of $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$, as indicated from EXAFS, can therefore also be supported by the

DRIFTS spectra which provide evidence for neighboring Pd centers to be linked by bridging nitrate coordination.

DRIFTS spectra of the precursor sample $\text{Pd}(\text{NH}_3)_4(\text{OH})_2$ or $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$ at room temperature show H–N–H vibrational modes of surface NH_3 species as bands at 1490 and 1345 cm^{-1} . Upon increasing temperature, the positions of the bands remain but also show steady decrease in intensity until complete removal by 225 °C. The detection of N_2O in the effluent gas by online mass spectrometry up to this temperature (Supporting Information, Figure S6b) shows that the NH_3 species are oxidized before leaving the chamber, consistent with previously reported catalytic activity for low temperature NH_3 oxidation at Pd surface. Following the removal of NH_3 , weak intensity absorption bands for N=O stretching vibrations of surface nitrosyl (NO^-_{ads}) are revealed at 1510 and 1556 cm^{-1} , respectively, highlighted by the red circle in Figure 5c. The formation of NO^- can be attributed to oxidative decomposition of NH_3 ligands, which then form strong interaction with the $\gamma\text{-Al}_2\text{O}_3$ surface. The nitrosyl absorption bands are present in the DRIFTS spectra up to 400 °C, which is consistent with previous temperature studies for stability of NO^- adsorbed onto the $\gamma\text{-Al}_2\text{O}_3$ surface.^{40–42} The M/Z 30 signal detected by online mass spectrometry of the effluent gas from sample $\text{Pd}(\text{NH}_3)_4(\text{OH})_2/\gamma\text{-Al}_2\text{O}_3$ between 390–450 °C (Supporting Information, Figure S6b) indicates a release of $\text{NO}^-_{(\text{ads})}$ as NO gas at this temperature. Note that in the case of the $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$ sample, the formation of NO^- species is unlikely as the

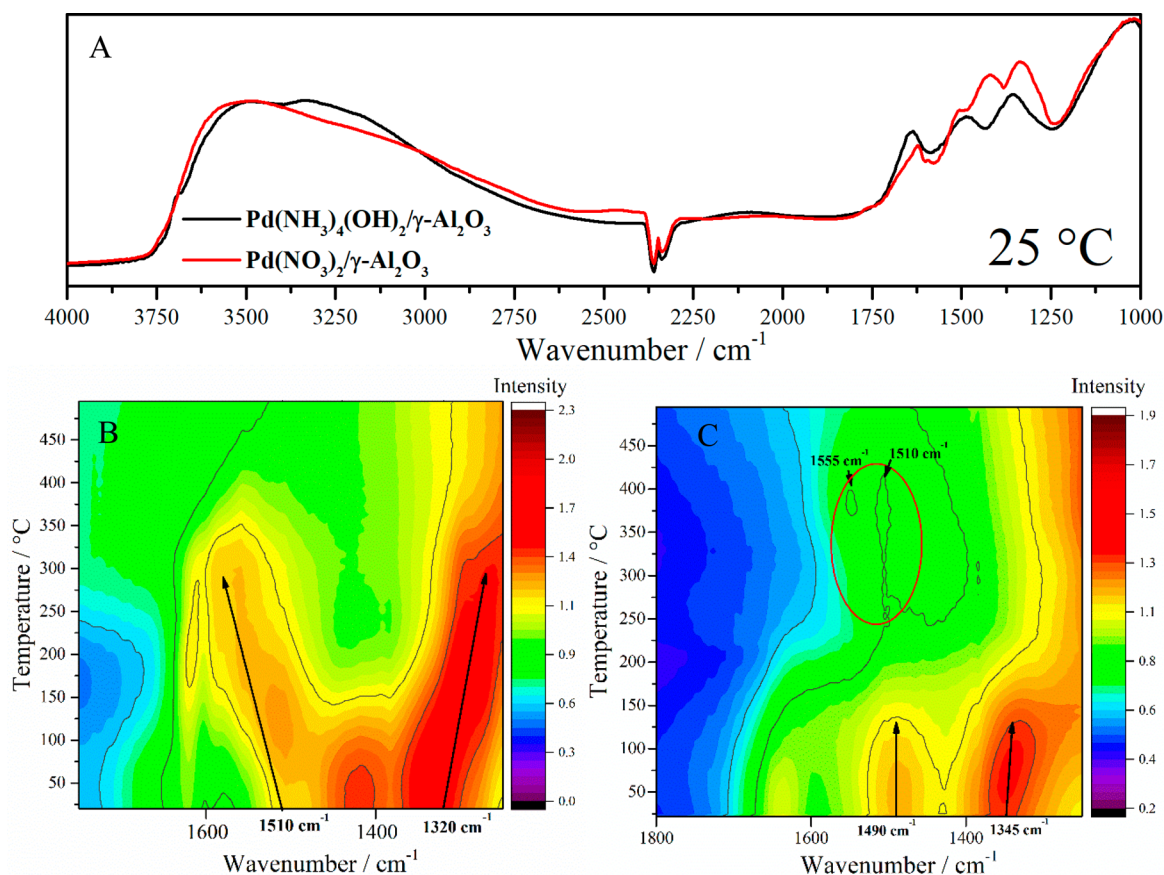


Figure 5. DRIFTS spectra of $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$ (red) and $\text{Pd}(\text{NH}_3)_4(\text{OH})_2/\gamma\text{-Al}_2\text{O}_3$ (black) at room temperature before calcination (A, top). Color maps of absorption intensities of time-resolved DRIFTS spectra at increasing temperature during calcination ramping period showing nitrate absorption bands of $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$ (B, bottom left) and $\text{Pd}(\text{NH}_3)_4(\text{OH})_2/\gamma\text{-Al}_2\text{O}_3$ (C, bottom right).

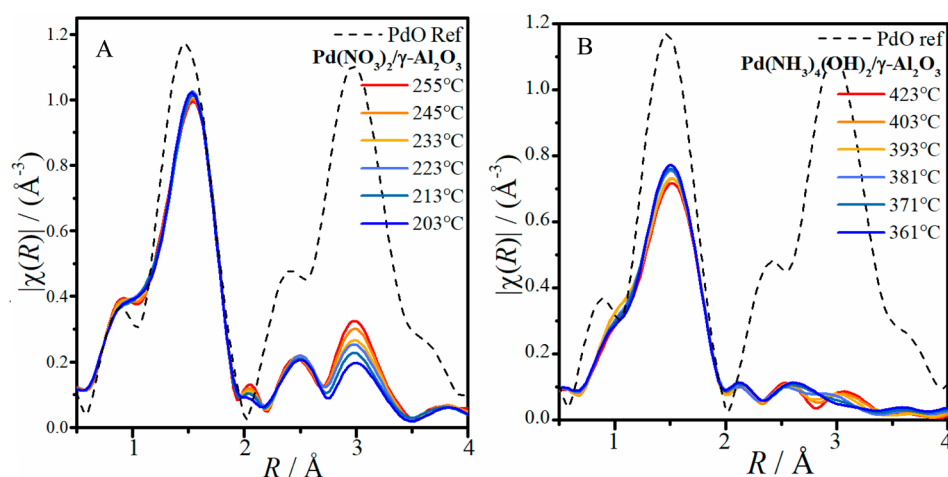


Figure 6. Nonphase corrected Fourier transformed EXAFS plots of (A) $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$ and (B) $\text{Pd}(\text{NH}_3)_4(\text{OH})_2/\gamma\text{-Al}_2\text{O}_3$ at increasing temperatures during the calcination ramp and reference PdO sample (dashed) at room temperature.

nitrogen of the nitrate ligands is already in higher oxidation state and thus is released in the oxidizing environment directly as NO gas (or NO_2), which is detected as M/Z 30 between 100 and 400 °C (Supporting Information, Figure S6a).

The nucleation and growth of PdO nanoparticles from each impregnated precursor can be followed from the amplitude of the scattering features of the Fourier transform Pd K-edge EXAFS data collected during the calcination ramp period, shown in Figure 6. The scattering features at 2.5 and 3 Å

correspond to scattering from nearest neighbor Pd atoms (located at real distances of 3.05 and 3.43 Å) in the tetrahedral PdO crystal structure.^{43,44} An increase in amplitude of these features at higher temperatures, despite increasing thermal motion and thus increased dampening from thermal disorder, indicates an increase in coordination to neighboring Pd atoms and thus growth of the PdO structure. For $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$, the scattering feature at 3 Å in the Fourier transform EXAFS data is first observed at room temperature (Figure 4d) and

starts to increase in amplitude when the sample is heated beyond 200 °C (Figure 6a). This shows that the distances between neighboring Pd atoms of the impregnated sample before calcination, thought to be associated via bridging nitrate coordination, do not change significantly upon adopting the Pd–O–Pd arrangement of tetrahedral PdO. For precursor sample $\text{Pd}(\text{NH}_3)_4(\text{OH})_2/\gamma\text{-Al}_2\text{O}_3$, the amplitude of scattering features in this low temperature period (200–360 °C) decreases with temperature due to the dampening effect of increasing EXAFS Debye–Waller factor (σ^2) with increasing thermal disorder. This is a clear indication that there is no increase in local coordination around the absorbing Pd atoms during this low temperature period. Introduction of the second shell Pd–Pd distance (3.43 Å) appears beyond 360 °C. This is shown in Figure 6b by the formation and growth in amplitude of the scattering feature at 3 Å. The coordination number of neighboring palladium atoms located at 3.43 Å from the absorbing palladium is plotted against temperature for both samples (Supporting Information, Figure S7). The analysis shows that nucleation of PdO nanoparticles from the $\text{Pd}(\text{NH}_3)_4(\text{OH})_2$ precursor occurs later in the calcination step than for the $\text{Pd}(\text{NO}_3)_2$ precursor, and growth occurs to a lesser extent. The PdO nanoparticle formation from ion exchanged molecular precursors $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$ and $\text{Pd}(\text{NH}_3)_4(\text{OH})_2/\gamma\text{-Al}_2\text{O}_3$ has been found to occur at differing temperatures of the calcination ramping period. $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$ showed significant intermolecular interaction before heat treatment, whereas $\text{Pd}(\text{NH}_3)_4(\text{OH})_2/\gamma\text{-Al}_2\text{O}_3$ showed isolation of Pd centers. This would be expected due to strong initial interactions of $[\text{Pd}(\text{NH}_3)_4]^{2+}$ cations with the negatively charged Al–O[−] surface sites (as alumina is deprotonated by OH[−] ions during incipient wetness impregnation) and is similarly observed for ion exchange of $\text{Pt}(\text{NH}_3)_4(\text{OH})_2$ onto SiO₂ surface.⁴⁵ It is worth noting that the temperature for NO[−] desorption from this sample, evidenced by loss of N=O stretching bands in DRIFTS spectra and presence of NO_(g) in effluent gas between 360–400 °C, correlates with the temperature at which Pd–Pd scattering paths were first identified from the EXAFS data. The stabilizing effect of NO species has been previously postulated in multiple studies by Sietsma et al., whereby introducing NO gas into the calcination atmosphere was shown to suppress nanoparticle agglomeration over metal oxide support.^{46,47} It is possible that a combination of both high initial Pd dispersion and NO[−]_(ads) stabilization contributed to the later formation and suppressed growth of PdO nanoparticles from the $\text{Pd}(\text{NH}_3)_4(\text{OH})_2$ precursor.

CONCLUSIONS

Through use of a combined XAFS/DRIFTS approach to study the formation of supported nanoparticles during calcination of impregnated metal oxides, a molecular insight into the role of different inorganic precursors was achieved. Pd/ $\gamma\text{-Al}_2\text{O}_3$ prepared from incipient wetness impregnation of aqueous $\text{Pd}(\text{NH}_3)_4(\text{OH})_2$ solution benefits from smaller Pd particle size and improved dispersion over the support, which in turn favors a lower light off temperature for CH₄ oxidation compared to that of the catalyst prepared from $\text{Pd}(\text{NO}_3)_2$.

The nucleation and growth of PdO nanoparticles during calcination was identified from the EXAFS data by an increase in the Pd–Pd scattering contribution located at a distance (3.43 Å) consistent with a Pd–Pd distance of crystalline PdO. For $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$, early association of Pd neighbors at this distance (3.43 Å) was observed in the EXAFS data of the

sample before heat treatment; the bridging nature of nitrate ligands in this sample, evidenced by N=O and O–N–O nitrate stretching frequencies in DRIFTS spectra, provides a rationale for the early assembly of Pd centers via bridging mechanism. Upon calcination of the $\text{Pd}(\text{NO}_3)_2/\gamma\text{-Al}_2\text{O}_3$ sample, there was a pronounced increase in the Pd–Pd scattering component at 3.43 Å between 200–250 °C, indicative of significant growth of PdO nanoparticles at this temperature. Conversely, the Pd centers from precursor $\text{Pd}(\text{NH}_3)_4(\text{OH})_2/\gamma\text{-Al}_2\text{O}_3$ appear to adopt isolated sites upon impregnation, evidenced by the absence of any Pd–Pd distances from the Pd EXAFS data, which can be explained by fast ion exchange of $[\text{Pd}(\text{NH}_3)_4]^{2+}$ with deprotonated Al–O[−] surface sites of the support. Although NH₃ ligands were shown to decompose from the impregnated sample below 200 °C, the formation of PdO nanoparticles (observed by Pd–Pd scattering from EXAFS data) did not occur until temperatures beyond 360 °C. Most NH₃ was partially oxidized and released as N₂O gas at 210 °C, but small amounts of NO[−] were captured by the $\gamma\text{-Al}_2\text{O}_3$ surface, evidenced by N=O stretching vibrations (1510 and 1556 cm^{−1}) in DRIFTS spectra and evolution of *M/Z* 30 in online mass spectrometry of the effluent gas upon their removal. Strong ionic interaction of $[\text{Pd}(\text{NH}_3)_2]^{2+}$ with the support and the possible role of adsorbed NO[−] result in stabilization of isolated PdO species on $\gamma\text{-Al}_2\text{O}_3$ for the preparation of a highly dispersed Pd/ $\gamma\text{-Al}_2\text{O}_3$ catalyst material.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/acs.chemmater.7b02552.

EXAFS fitting parameters, Raman emission spectra, XANES spectra, EXAFS particle size calculations, catalytic light off curves, reference DRIFTS spectra, TGA curves, and mass spectrometry data (PDF)

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Notes

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ABBREVIATIONS

DRIFTS, diffuse reflectance infrared Fourier transform spectroscopy
EXAFS, extended X-ray absorption fine structure
XANES, X-ray absorption near-edge structure
TEM, transmission electron microscopy
FTIR, Fourier transformed infrared
XAFS, X-ray absorption fine structure
XRD, X-ray diffraction

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